1. **Charge-transfer polarization** limits current because the rate of electron transfer at the electrode is slow. **Concentration polarization** limits the current because mass transfer of the electroactive species from the bulk solution to the film around the electrode is slow.

2. Mass transport occurs by: diffusion because of a concentration gradient. electrostatic attraction. mechanical convection (stirring).

3. Concentration polarization is decreased with: higher reactant concentration. lower total electrolyte concentration. more stirring. larger electrode surface area.

4. Charge-transfer polarization is more likely with: higher current densities. lower temperature. softer metal electrodes. reactions with gaseous products.

5. **Faradaic** processes involve electron-transfer (redox) reactions. **Non-faradaic** processes involve current flow in the absence of redox reactions (e.g., charging current, ion migration).